

Spectator ion-

Net ionic equation-

Using a solubility table salt of alkali metals are _____, ammonia is also soluble. Nitrates, _____, sulphates and chlorides(except for those attached to Pb^{2+} , Ag^+ , Ba^{2+} , Sr^{2+} , Ca^{2+} . Insoluble compounds usually contain _____, _____, _____, _____ and hydroxides. (table 11.3)

Water is polar because of _____.

There are _____ valence orbitals in the valence level of representative elements.

Surface tension-

Vapor pressure-

Solute-

Solvent-

Solvation-

Electrolyte-

Non-electrolyte-

Hydrate-

#valence electrons=_____

A maximum of _____ electrons can occupy each individual orbital in the valence orbitals.

Cation-

Anion-

Octet rule:

Lone pair:
(or non-bonding pair)

Bonding pair:

Draw Lewis electron dot diagrams below for the following elements: H, Mg, I, S, P, C, B

H•						
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Ionic bond-

Chemical formula-

Formula Unit-

Three properties of ionic compounds are-

1-

2-

3-

Metallic bonds-

An alloy is-

What 3 factors involve the contact of a solute and its solvent-

1- _____ 2- _____ 3- _____

Saturated means- _____ Unsaturated- _____

Units for solubility- _____ Solubility is affected by _____ and pressure

The element with the highest electronegativity is _____.

Molarity- give the units and a formula→ _____

Give a formula for working with dilution of a solution: $V_1C_1 =$ _____

%by volume= _____

%by mass= _____

Calculate the volume of concentrated phosphoric acid(14.6 mol/L) that must be diluted to prepare 500 mL of a 1.25 mol/L solution.

Calculate the mass of $\text{KHC}_4\text{H}_4\text{O}_6$ that is measured to prepare 100 mL of a 0.150mol/L standard solution.

Calculate the concentration of a solution containing 16.0 g of sodium hydroxide in 2.0 L of water.

Formula	Lewis electron dot diagram	Structural diagram		Polarity
Ammonia				
Fluorine				
Acetic acid				
water				
Ethanol				
propane				
Boron trichloride				

Give 2 examples of elements that have the following number of bonds:

Single:	Double:	Triple:	All 3 types:
ex. _____	ex. _____	ex. _____	ex. -N= , -N- , N _____

Do p 343 #29, p 344 #, 32a,33a,34all,35a,c, p 193 7a,c,9all,p 481 #9, p 482#11,
p 484 #12p 485 14,15

